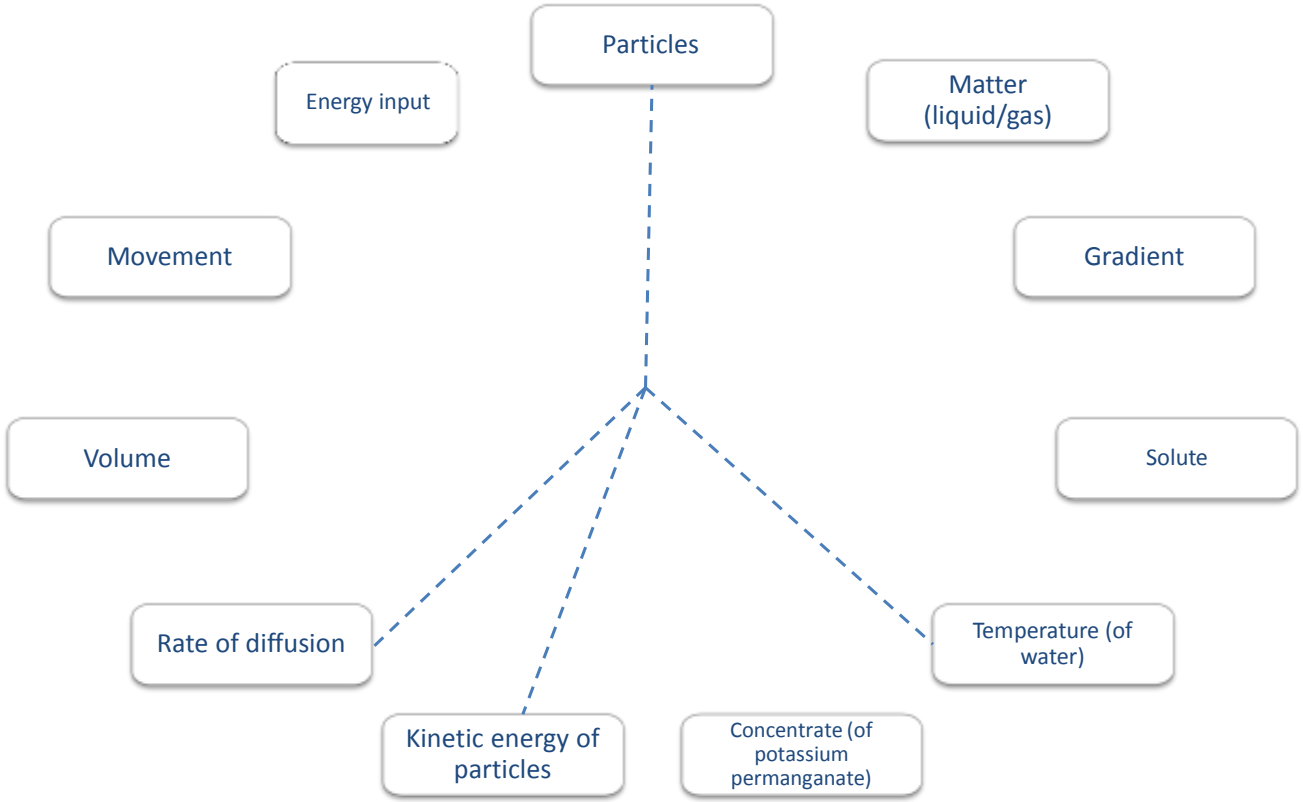
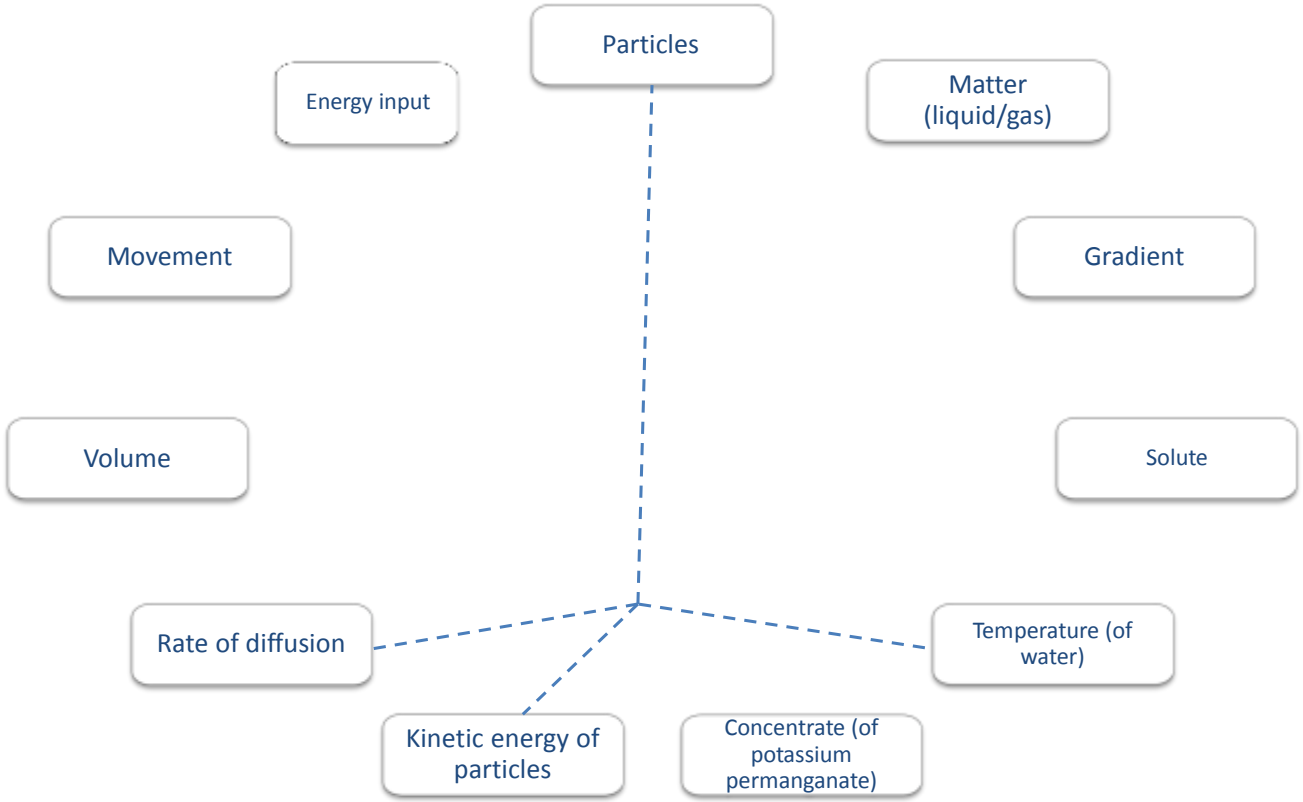
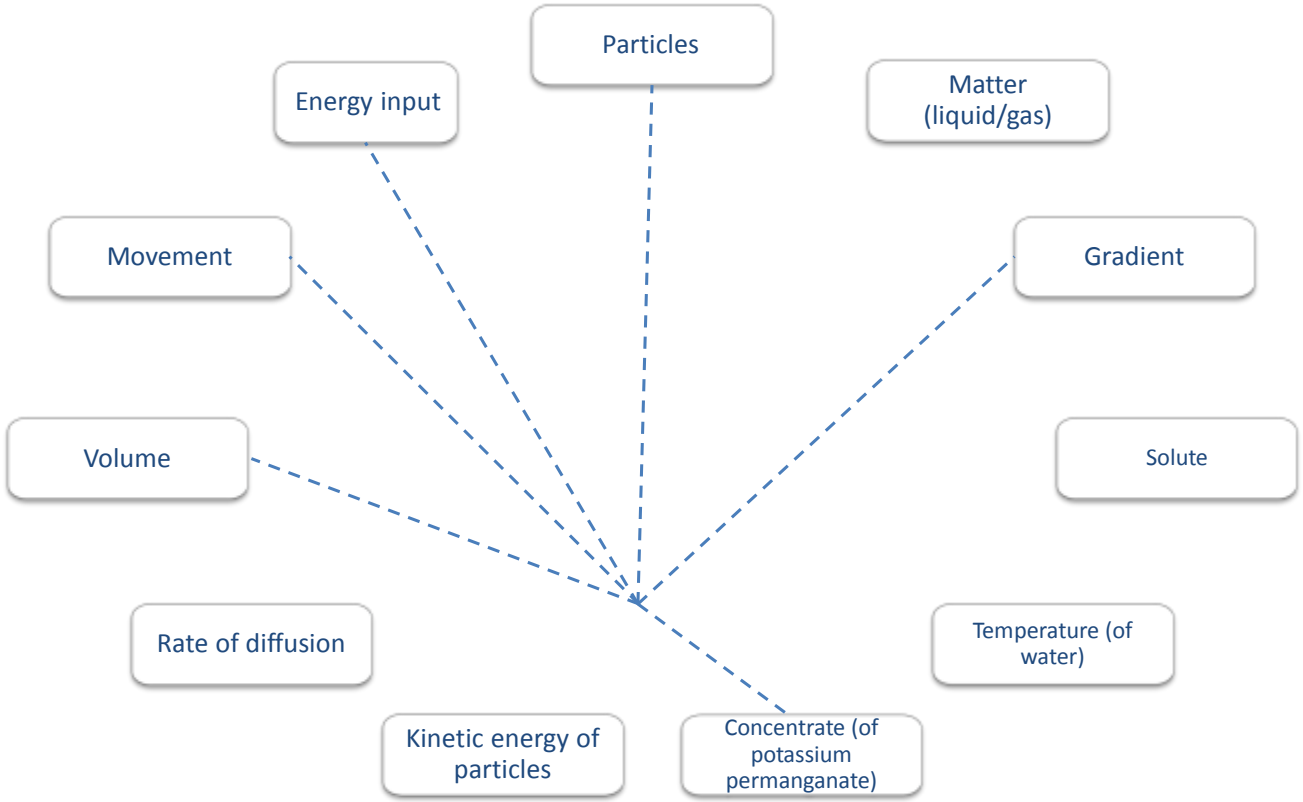
Diffusion conception 1: The rate of diffusion is higher in gases as particles in a gas can move faster.



Diffusion conception 2: An increase in temperature leads to an increase in the particles' kinetic energy. The higher the particles' kinetic energy the faster the rate of diffusion.



Diffusion conceptions 3: Diffusion is the movement of particles from a volume of high concentration to lower concentration, down a gradient without further energy input.



What is to be discerned	**Critical aspects**				
	Initial concentration (of potassium permanganate)	Temperature	Volume (of water)	Rate of diffusion	Volume (of potassium permanganate)
Increasing temperature increases rate of diffusion	i	v	i	v	i
Increase in volume (of water) lowers rate of diffusion	i	i	v	v	v
Increased initial concentration (of potassium permanganate) increases rate of diffusion	v	i	i	v	v